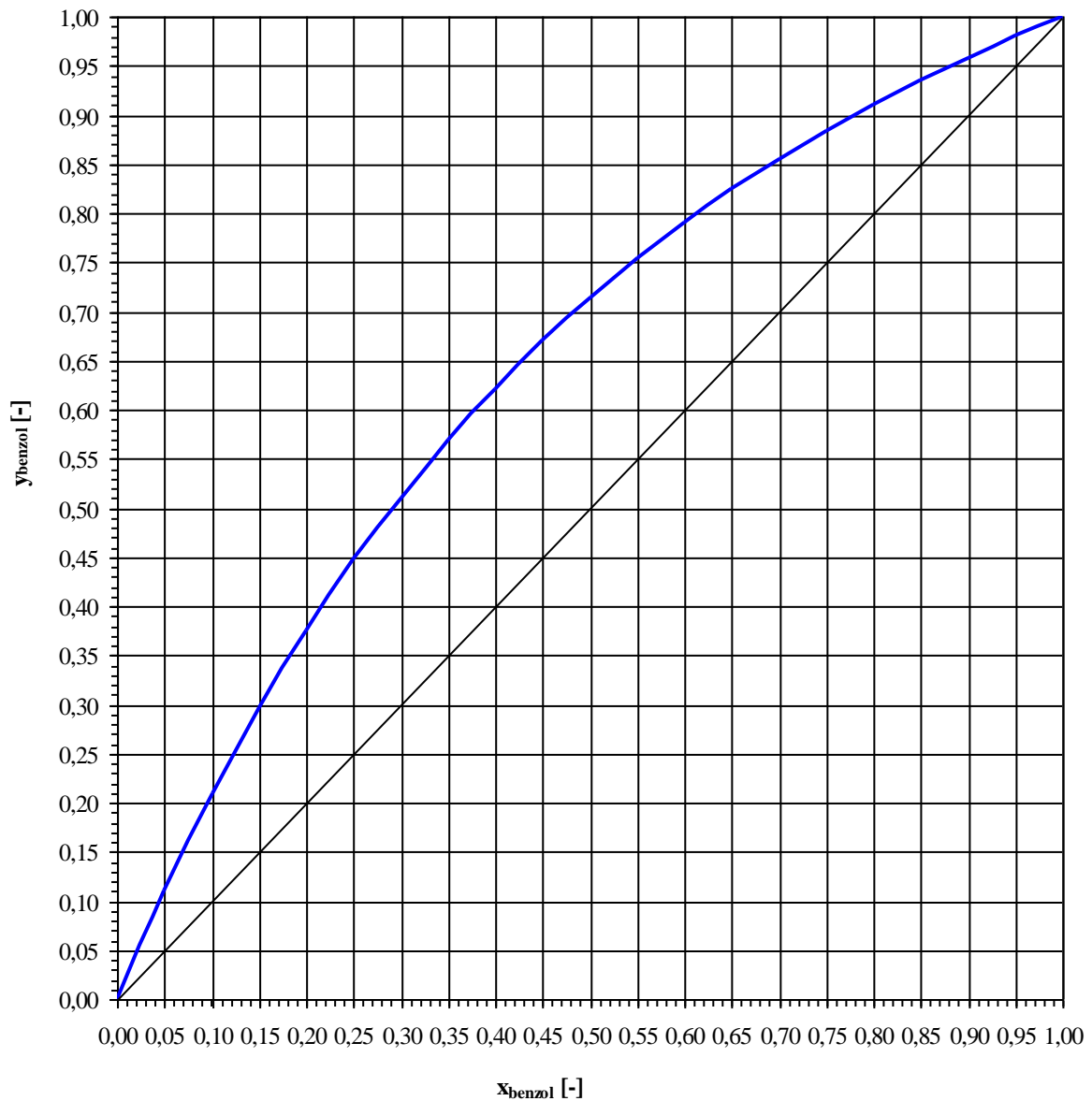


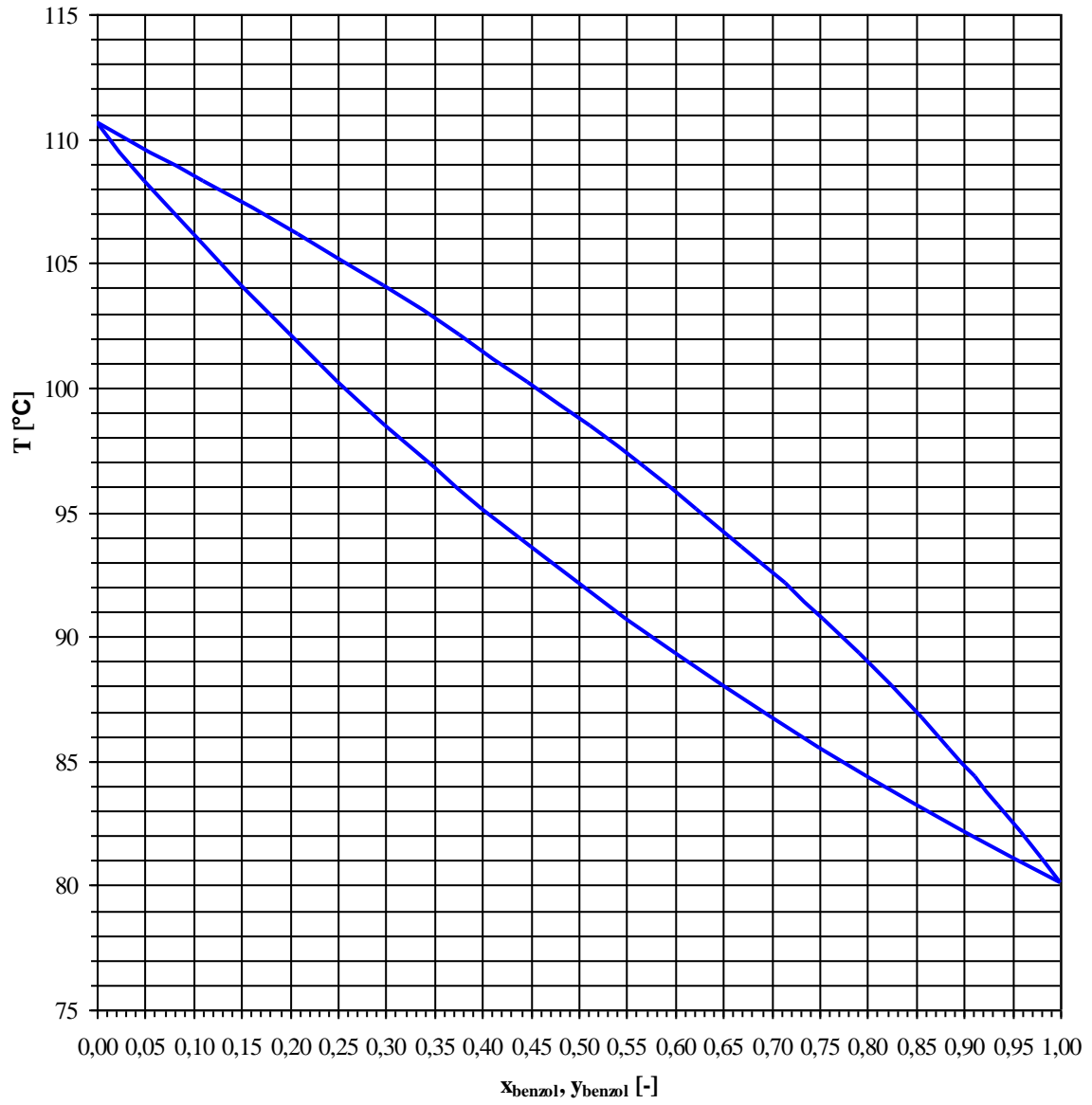
## Elegyek adatai

### *Benzol–toluol*

		<b>Benzol</b>	<b>Toluol</b>
<b>Antoine-konstansok</b>	<i>A</i>	6,90565	6,95464
	<i>B</i>	1211,033	1344,80
	<i>C</i>	220,79	219,482
<b><i>M</i> [g/mol]</b>		78	92
<b><i>T<sub>p</sub></i> [°C] (<i>p</i> = 1 atm)</b>		80,1	110,6



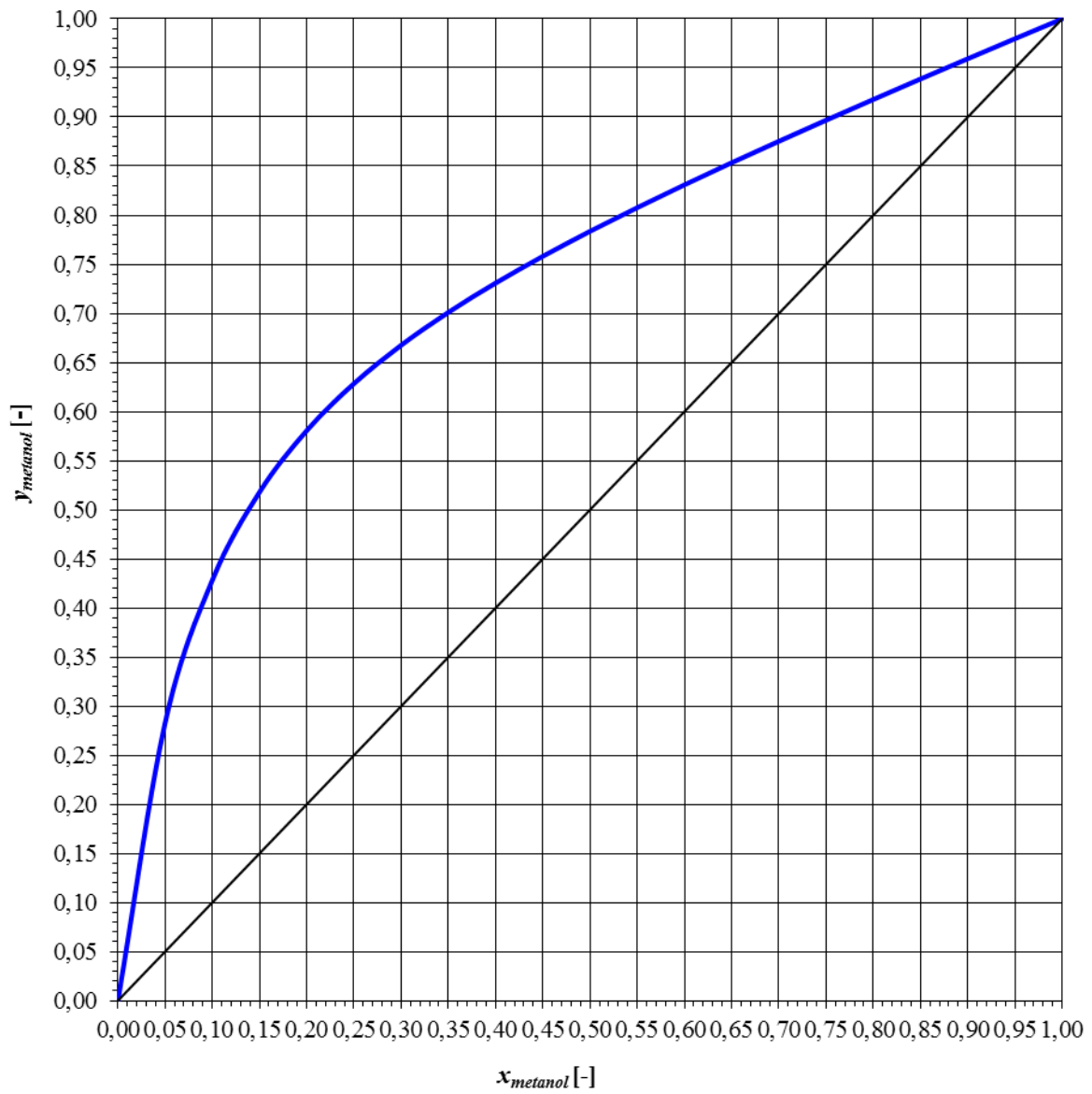
*X.1. ábra Benzol–toluol elegy egyensúlyi diagramja (*p* = 1 atm)*



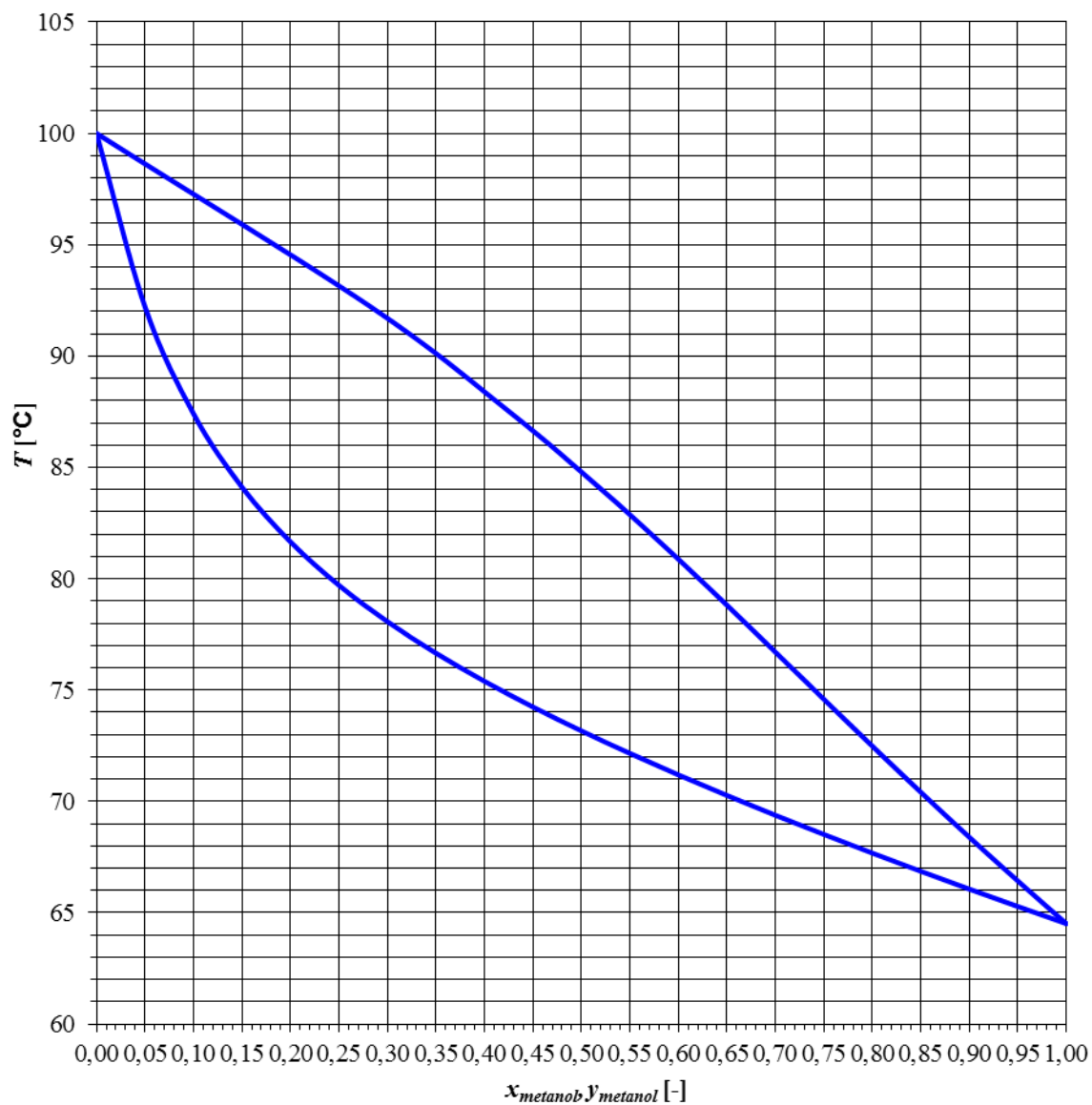
X.2. ábra Benzol–toluol elegy forrpont diagramja ( $p = 1 \text{ atm}$ )

*Metanol-víz*

	Metanol	Víz
$M$ [g/mol]	32	18
$T_{fp}$ [°C] ( $p = 1$ atm)	64,5	100,0



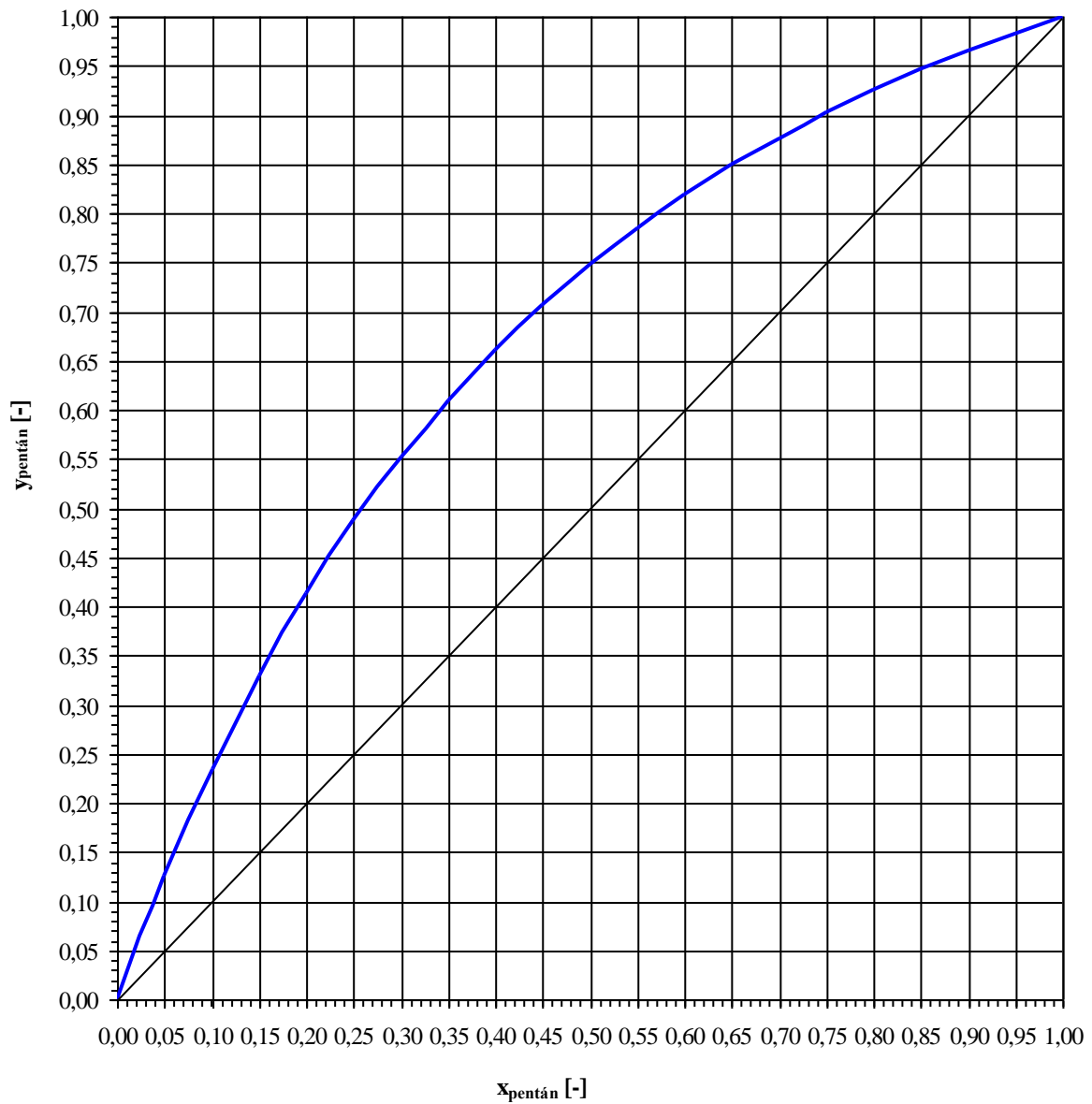
*X.3. ábra Metanol-víz elegy egyensúlyi diagramja ( $p = 1$  atm)*



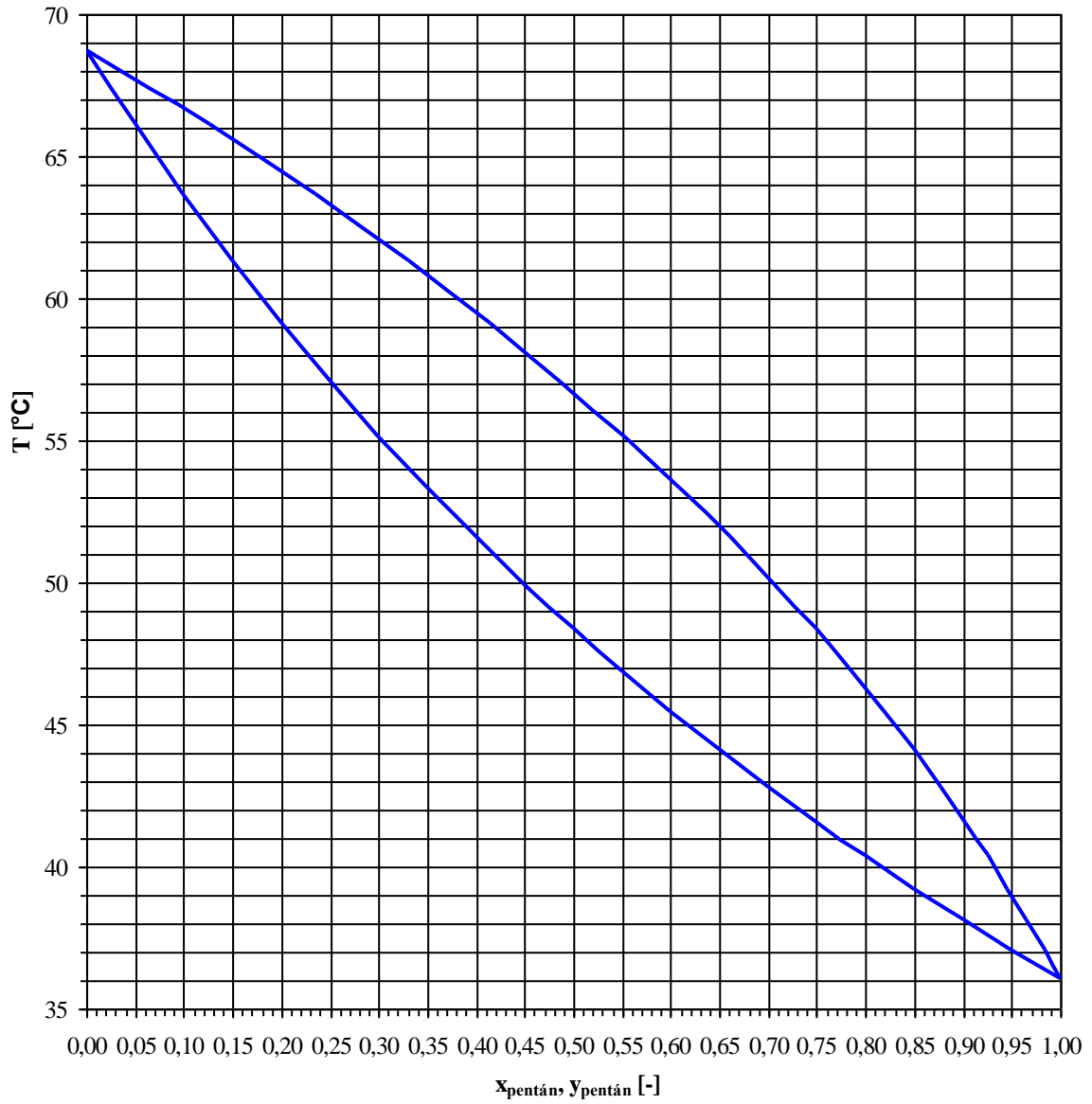
X.4. ábra Metanol-víz elegy forrpont diagramja ( $p = 1 \text{ atm}$ )

**Pentán–hexán**

		<b>Pentán</b>	<b>Hexán</b>
<b>Antoine-konstansok</b>	<b>A</b>	6,87632	6,87024
	<b>B</b>	1075,78	1168,72
	<b>C</b>	233,205	224,21
<b>M [g/mol]</b>		72	86
<b>T<sub>p</sub> [°C] (p = 1 atm)</b>		36,0	68,7



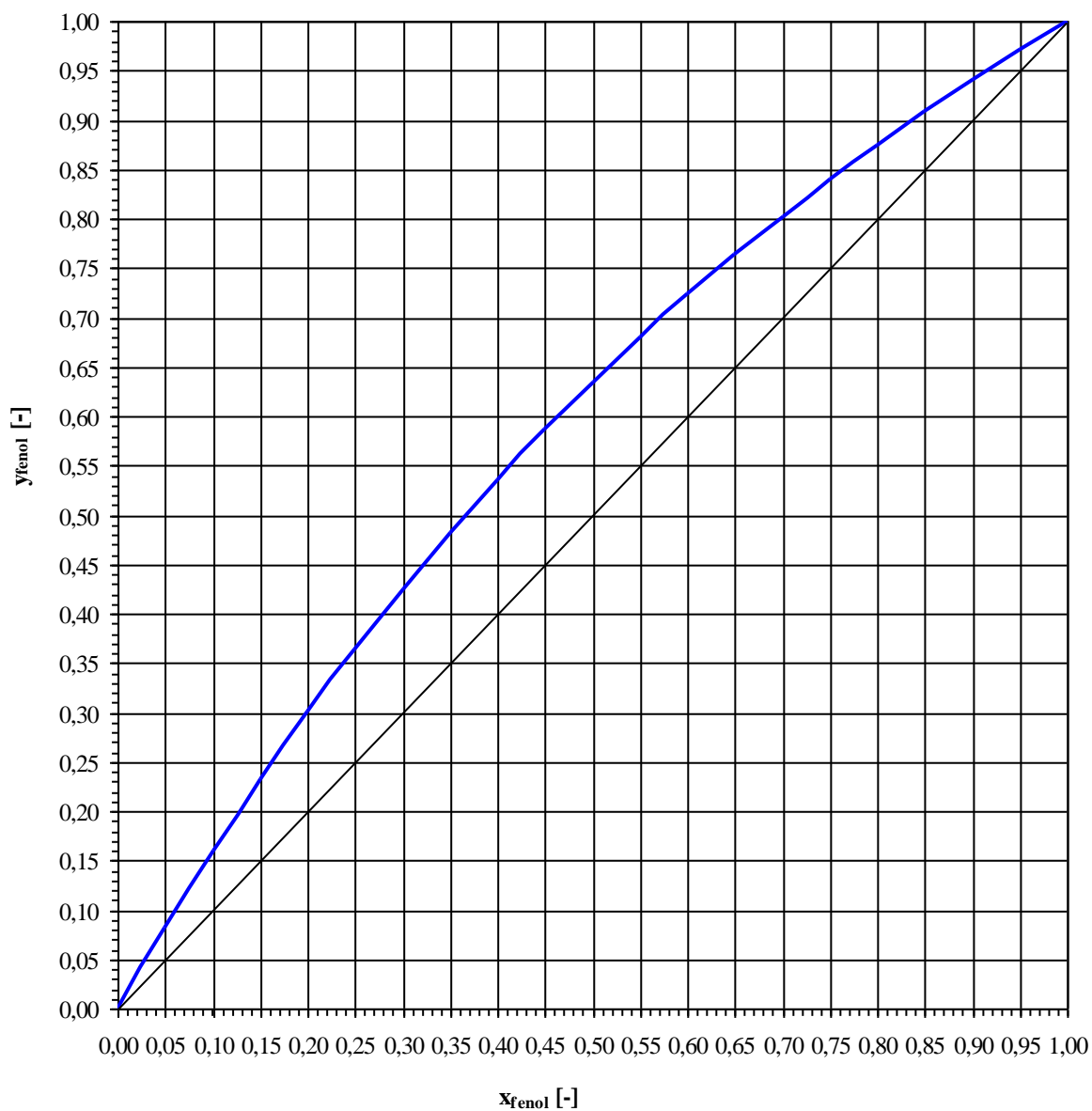
X.5. ábra Pentán–hexán elegy egyensúlyi diagramja (p = 1 atm)



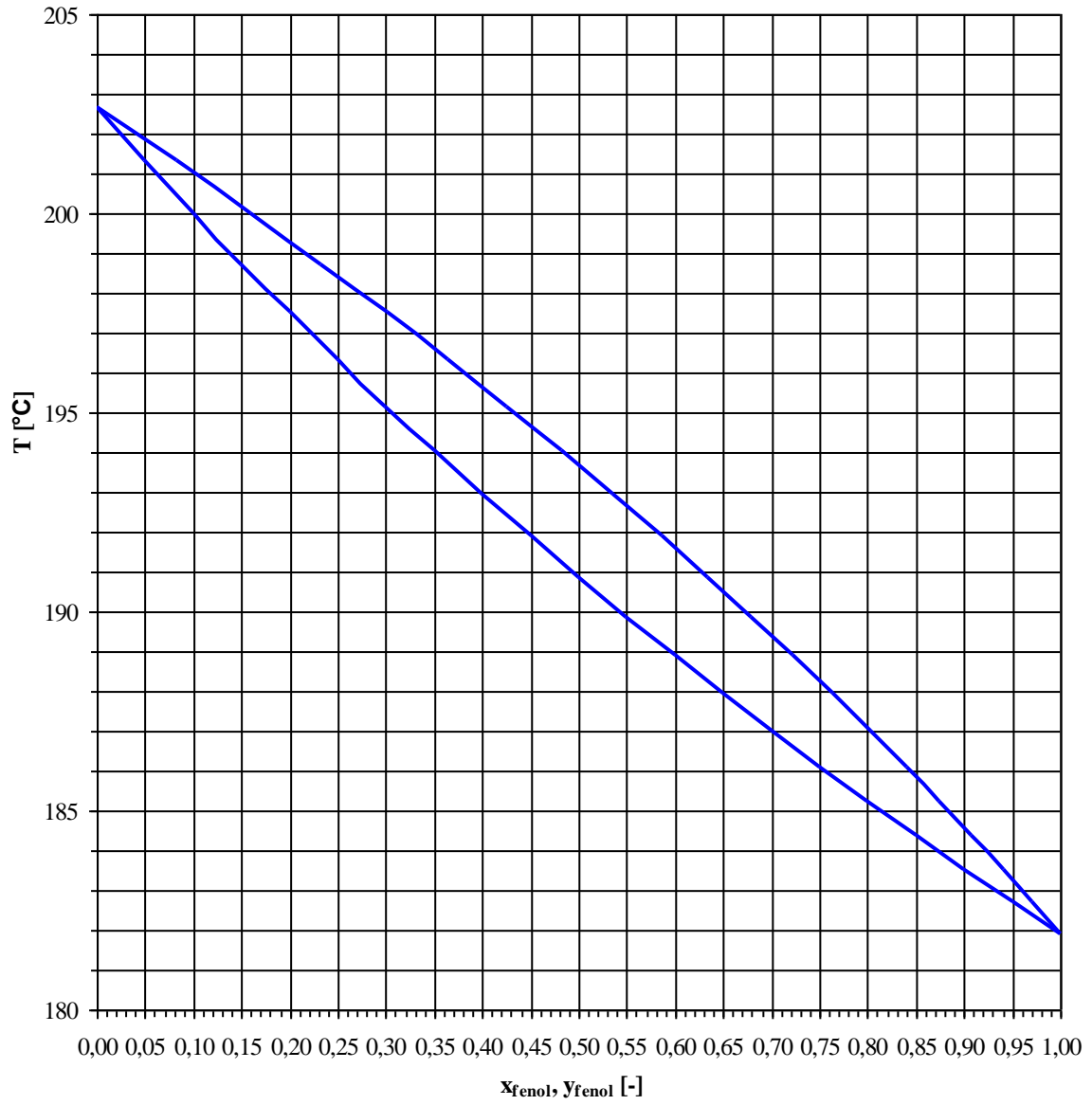
X.6. ábra Pentán–hexán elegy forrpont diagramja ( $p = 1 \text{ atm}$ )

*Fenol–metakrezol*

		Fenol	Metakrezol (3-metilfenol)
Antoine- konstansok	A	6,93051	6,76147
	B	1382,65	1355,92
	C	159,493	146,73
$M$ [g/mol]		94	108
$T_p$ [°C] ( $p = 1$ atm)		181,9	202,7



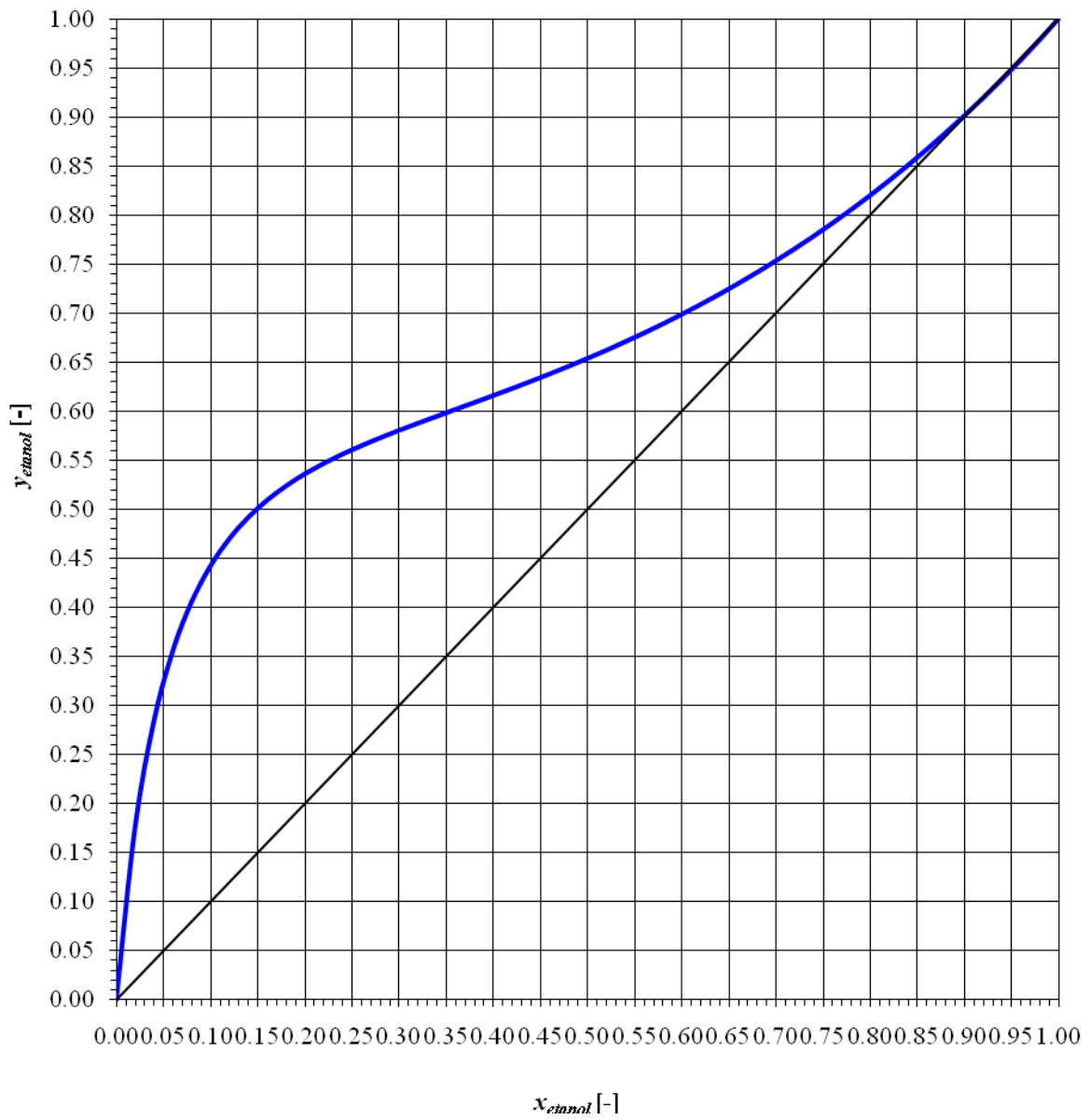
*X.7. ábra Fenol–metakrezol elegy egyensúlyi diagramja ( $p = 1$  atm)*



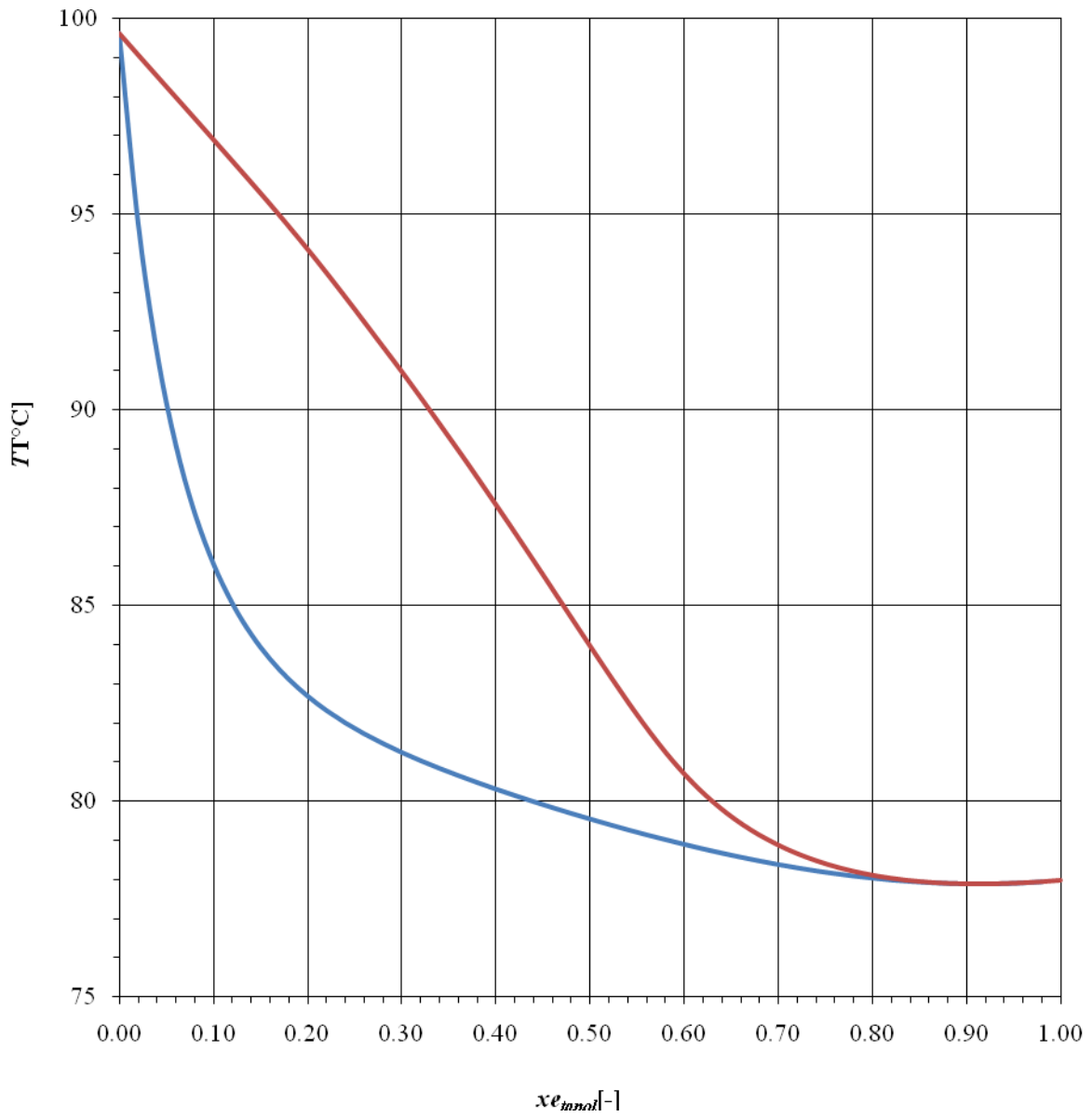
X.8. ábra: Fenol–metakrezol elegy forrpont diagramja ( $p = 1 \text{ atm}$ )



*Etanol – víz*



*x.9 ábra: Etanol - víz elegy egyensúlyi diagramja ( $p = 1 \text{ atm}$ )*



x.10 ábra: Etanol - víz elegy forrpont diagramja ( $p = 1 \text{ atm}$ )